

## The Crystal Structures of the Tetragonal and Monoclinic Forms of Bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium

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Bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium crystallizes from ethanol solution in two crystalline forms, tetragonal and monoclinic. The monoclinic form is isomorphous with bis-(*N*-isopropyl-3-methylsalicylaldiminato)nickel. The crystal structures of both forms have been determined from three-dimensional X-ray diffraction data. The cell dimensions for the tetragonal form are:  $a_0 = 18.178$ ,  $c_0 = 12.879$  Å; and for the monoclinic form are:  $a_0 = 11.297$ ,  $b_0 = 9.755$ ,  $c_0 = 9.844$  Å,  $\beta = 107^\circ 51'$ . The space group for the tetragonal form is  $I4_1/a$  with eight molecules per cell and for the monoclinic form is  $P2_1/c$  with two molecules per cell. The palladium atoms have trans-planar coordination in both cases. The bond distances and angles in the chelate are similar to those reported for other salicylaldimine complexes, but differences in packing cause significant bending of the nearly planar molecules.

### Introduction

In an attempt to understand the structural reasons for the variations in coordination configuration in a series of substituted salicylaldimine chelates of nickel(II), we have undertaken the investigation of the crystal structures of a number of salicylaldimine chelates (Fox, Orioli, Lingafelter & Sacconi, 1964; Braun & Lingafelter, 1966, 1967*a, b*). The palladium(II) ion has been selected as an ion whose coordination configuration would probably be planar in all cases (Sacconi, Ciampolini, Maggio & Del Re, 1960). We have now completed the structural investigation of the two forms of bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium.

### Experimental

Bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium was prepared by the method of Sacconi, Paoletti & Del Re (1957), using 3-methylsalicylaldehyde which was prepared by the method of Duff (1941). Purification and recrystallization were carried out from ethyl alcohol by slow evaporation. The crystals are yellow in color and two forms of crystals are easily separable under the microscope. The crystals of the tetragonal form are rectangular with pyramidal ends while those of the monoclinic form are rhomb-shaped plates. Well-defined crystals of the two forms were chosen for study. They had approximate dimensions  $0.09 \times 0.10 \times 0.12$  mm and  $0.14 \times 0.15 \times 0.03$  mm, respectively. All measurements were made with Mo  $K\alpha$  radiation ( $\lambda = 0.71069$  Å) on a Picker Single Crystal Diffractometer equipped with pulse-height analyser and a General Electric Single Crystal Orienter (Furnas, 1957). The cell constants and their estimated standard deviations were obtained by least-squares fit of 21  $\theta$  values in the case of the tetragonal form and of 10  $\theta$  values in the case of the monoclinic form. The final values are: tetragonal,  $a_0 = 18.178(1)$ ,  $c_0 = 12.879(2)$  Å; monoclinic,

$a_0 = 11.297(1)$ ,  $b_0 = 9.755(1)$ ,  $c_0 = 9.844(1)$  Å,  $\beta = 107^\circ 51'(1)$ . The systematic absence of  $hkl$  for  $h+k+l \neq 2n$ ,  $00l$  for  $l \neq 4n$ , and  $hk0$  for  $h \neq 2n$  and  $k \neq 2n$  identified the space group  $I4_1/a$  for the tetragonal form, while the systematic absence of  $0k0$  for  $k$  odd and  $h0l$  for  $l$  odd identified the space group  $P2_1/c$  for the monoclinic form. The crystal densities were found to be  $1.40$  g.cm $^{-3}$  and  $1.45$  g.cm $^{-3}$ , respectively. The calculated densities, assuming eight molecules in the tetragonal cell and two molecules in the monoclinic cell, are  $1.429$  g.cm $^{-3}$  and  $1.473$  g.cm $^{-3}$ , respectively. The linear absorption coefficient for Mo  $K\alpha$  radiation is  $8.76$  cm $^{-1}$ .

All calculations were carried out on an IBM 7094 computer with the programs written or adapted by Stewart (1964). Diffractometer settings for each of the possible diffraction maxima were calculated for the  $\omega$ - $2\theta$  scan method using the recommended formula of Alexander & Smith (1964): scan range =  $A + B \tan \theta$ , with  $A$  set equal to 1.4 and  $B$  equal to 1.00.

At the beginning and at the end of each six-hour interval of data collection, the intensities of four standard reflections were measured. The reciprocal of the mean value of these eight intensity measurements was then used as a relative scale factor for all intensities measured during that interval. These scale factors showed no systematic variation with time and the maximum variation from the scale factor for the initial interval was 0.5% for the tetragonal crystal data and 3.5% for the monoclinic crystal data.

Intensities were measured for 1847 reflections, lying in the range  $0 < 2\theta < 50^\circ$ , for the tetragonal crystal and 1284 reflections, lying in the range  $0 < 2\theta < 45^\circ$ , for the monoclinic crystal. 563 of the reflections from the tetragonal crystal and 495 of the reflections from the monoclinic crystal gave net intensities less than twice their standard deviations (calculated from counting statistics) and were assigned intensity values of  $2\sigma_I$  and treated as 'unobserved'.

Lorentz and polarization factors were applied and relative structure factors were calculated. No corrections were made for absorption or for anomalous dispersion.

### Structure determination

#### Tetragonal form

From a survey of the intensity distribution in the case of the tetragonal form it may be seen that, in general, only reflections with  $h+k+l=4n$  and  $h, k$  and  $l$  all even or with  $h+k$  and  $l$  both odd are strong. This suggests that the palladium atoms lie in positions (c) or (d) (*International Tables for X-ray Crystallography*, 1965), and (c) can be chosen without loss of generality. For all calculations the origin was chosen at a center of symmetry.

An initial structure factor calculation, carried out by placing the palladium atom of the asymmetric unit at the origin, gave  $R$  equal to 0.31.  $R$  is defined throughout as  $\Sigma ||F_o| - |F_c|| / \Sigma |F_o|$ , with summation over the unique, observed reflections only. The atomic scattering factors used were those of Thomas & Umeda (1957) for palladium, those of Berghuis, Haanappel, Potters, Loopstra, MacGillavry & Veenendaal (1955) for oxygen and carbon, and those from Table 2 of Stewart, Davidson & Simpson (1965) for hydrogen.

A three-dimensional Fourier synthesis, using phases determined by palladium alone, was calculated. On the

Fourier map the nitrogen atom was located easily. Another structure factor calculation based on palladium and nitrogen brought  $R$  to 0.28. A Fourier synthesis phased on these atoms revealed the positions of the oxygen atom and carbon atoms C(1), C(2) and C(3). Another three-dimensional Fourier synthesis phased on these atoms enabled the positions of all the remaining non-hydrogen atoms to be located. During all stages of these structure factor and Fourier calculations an overall isotropic temperature factor,  $B=2.72$ , was used.

Further refinement was carried out with the full-matrix least-squares program of Busing & Levy (1959). The function minimized was  $\Sigma w(|F_o| - |F_c|)^2$ . Four cycles of least-squares refinement with individual isotropic temperature factors and with all  $w=1$  (except that, in each cycle, unobserved reflections with  $F_c < F_o$  were given  $w=0$ ) reduced  $R$  to 0.063. Three more cycles of least-squares refinement with anisotropic temperature factors reduced  $R$  to 0.049.

At this stage, a  $\Delta F$  synthesis showed the positions of all hydrogen atoms. The hydrogen atom coordinates thus obtained were refined by three cycles of least-squares refinement, treating the non-hydrogen atoms as fixed.

A further series of four cycles of combined refinement of both non-hydrogen and hydrogen atoms with mixed temperature factors (individual anisotropic tem-

Table 1. Tetragonal form

Positional parameters ( $\times 10^4$ ) and thermal parameters ( $\times 10^2$ )

The thermal parameters are of the form  $\exp(-\frac{1}{4} \sum_{i=1}^3 \sum_{j=1}^3 B_{ij} h_i h_j a_i^* a_j^*)$ .

	$x/a$	$y/b$	$z/c$	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
Pd	2500	2500	2500	279 (2)	273 (2)	303 (2)	28 (2)	12 (2)	26 (2)
O	2927 (1)	2557 (2)	1081 (2)	321 (13)	356 (14)	298 (14)	104 (12)	58 (11)	60 (11)
N	1821 (2)	3362 (2)	2197 (3)	311 (16)	274 (15)	327 (19)	104 (12)	9 (14)	26 (13)
C(1)	2096 (2)	3453 (3)	0343 (3)	346 (20)	271 (18)	338 (21)	40 (15)	1 (16)	-16 (16)
C(2)	2687 (2)	2940 (2)	0291 (3)	315 (19)	263 (18)	306 (21)	-30 (15)	-9 (14)	29 (14)
C(3)	3041 (2)	2838 (2)	-0673 (3)	392 (22)	304 (21)	398 (24)	-21 (17)	74 (19)	14 (18)
C(4)	2798 (3)	3205 (3)	-1530 (4)	617 (31)	521 (28)	268 (24)	28 (23)	65 (22)	65 (21)
C(5)	2207 (3)	3696 (3)	-1494 (4)	653 (33)	535 (30)	383 (27)	140 (25)	-8 (26)	121 (24)
C(6)	1871 (3)	3810 (3)	-0567 (4)	508 (27)	415 (25)	311 (24)	89 (20)	-26 (20)	63 (19)
C(7)	1731 (2)	3631 (2)	1279 (3)	337 (21)	274 (20)	398 (24)	107 (17)	-39 (18)	52 (18)
C(8)	1380 (3)	3700 (3)	3046 (4)	453 (26)	435 (24)	311 (23)	169 (20)	103 (20)	56 (42)
C(9)	1452 (5)	4527 (4)	3072 (6)	1009 (56)	496 (32)	467 (36)	245 (34)	131 (32)	-054 (28)
C(10)	0606 (5)	3419 (7)	2969 (9)	526 (41)	1163 (73)	772 (59)	11 (41)	252 (41)	-144 (50)
C(11)	3682 (3)	2313 (4)	-0732 (5)	505 (30)	567 (34)	470 (31)	66 (25)	207 (26)	94 (26)

Positional parameters ( $\times 10^3$ ) and isotropic thermal parameters

	$B$			
H(4)	296 (2)	312 (2)	-213 (3)	2.7 (1)
H(5)	203 (3)	396 (3)	-203 (4)	5.8 (1)
H(6)	147 (3)	416 (3)	-049 (4)	4.8 (1)
H(7)	138 (2)	399 (2)	122 (3)	2.5 (1)
H(8)	157 (2)	351 (2)	364 (3)	2.3 (1)
H(91)	127 (4)	470 (3)	372 (5)	9.0 (2)
H(92)	206 (4)	470 (4)	306 (6)	10.4 (2)
H(93)	125 (4)	472 (4)	253 (6)	8.6 (2)
H(101)	043 (3)	354 (3)	335 (4)	3.7 (2)
H(102)	032 (4)	351 (4)	226 (6)	9.9 (3)
H(103)	051 (4)	288 (5)	274 (8)	14.5 (3)
H(111)	391 (3)	235 (3)	-136 (4)	4.7 (1)
H(112)	415 (3)	249 (3)	-019 (5)	6.3 (1)
H(113)	355 (2)	184 (2)	-056 (4)	3.7 (1)

Table 2. Observed and calculated structure factors for the tetragonal form  
Columns are  $h$ ,  $10F_o$  and  $10F_c$ . Unobserved reflections are marked with \*.

$h$	$10F_o$	$10F_c$	$h$	$10F_o$	$10F_c$	$h$	$10F_o$	$10F_c$	$h$	$10F_o$	$10F_c$	$h$	$10F_o$	$10F_c$	$h$	$10F_o$	$10F_c$
000	1000	1000	100	1000	1000	200	1000	1000	300	1000	1000	400	1000	1000	500	1000	1000
100	1000	1000	110	1000	1000	210	1000	1000	310	1000	1000	410	1000	1000	510	1000	1000
200	1000	1000	120	1000	1000	220	1000	1000	320	1000	1000	420	1000	1000	520	1000	1000
300	1000	1000	130	1000	1000	230	1000	1000	330	1000	1000	430	1000	1000	530	1000	1000
400	1000	1000	140	1000	1000	240	1000	1000	340	1000	1000	440	1000	1000	540	1000	1000
500	1000	1000	150	1000	1000	250	1000	1000	350	1000	1000	450	1000	1000	550	1000	1000
600	1000	1000	160	1000	1000	260	1000	1000	360	1000	1000	460	1000	1000	560	1000	1000
700	1000	1000	170	1000	1000	270	1000	1000	370	1000	1000	470	1000	1000	570	1000	1000
800	1000	1000	180	1000	1000	280	1000	1000	380	1000	1000	480	1000	1000	580	1000	1000
900	1000	1000	190	1000	1000	290	1000	1000	390	1000	1000	490	1000	1000	590	1000	1000
1000	1000	1000	200	1000	1000	300	1000	1000	400	1000	1000	500	1000	1000	600	1000	1000
1100	1000	1000	210	1000	1000	310	1000	1000	410	1000	1000	510	1000	1000	610	1000	1000
1200	1000	1000	220	1000	1000	320	1000	1000	420	1000	1000	520	1000	1000	620	1000	1000
1300	1000	1000	230	1000	1000	330	1000	1000	430	1000	1000	530	1000	1000	630	1000	1000
1400	1000	1000	240	1000	1000	340	1000	1000	440	1000	1000	540	1000	1000	640	1000	1000
1500	1000	1000	250	1000	1000	350	1000	1000	450	1000	1000	550	1000	1000	650	1000	1000
1600	1000	1000	260	1000	1000	360	1000	1000	460	1000	1000	560	1000	1000	660	1000	1000
1700	1000	1000	270	1000	1000	370	1000	1000	470	1000	1000	570	1000	1000	670	1000	1000
1800	1000	1000	280	1000	1000	380	1000	1000	480	1000	1000	580	1000	1000	680	1000	1000
1900	1000	1000	290	1000	1000	390	1000	1000	490	1000	1000	590	1000	1000	690	1000	1000
2000	1000	1000	300	1000	1000	400	1000	1000	500	1000	1000	600	1000	1000	700	1000	1000
2100	1000	1000	310	1000	1000	410	1000	1000	510	1000	1000	610	1000	1000	710	1000	1000
2200	1000	1000	320	1000	1000	420	1000	1000	520	1000	1000	620	1000	1000	720	1000	1000
2300	1000	1000	330	1000	1000	430	1000	1000	530	1000	1000	630	1000	1000	730	1000	1000
2400	1000	1000	340	1000	1000	440	1000	1000	540	1000	1000	640	1000	1000	740	1000	1000
2500	1000	1000	350	1000	1000	450	1000	1000	550	1000	1000	650	1000	1000	750	1000	1000
2600	1000	1000	360	1000	1000	460	1000	1000	560	1000	1000	660	1000	1000	760	1000	1000
2700	1000	1000	370	1000	1000	470	1000	1000	570	1000	1000	670	1000	1000	770	1000	1000
2800	1000	1000	380	1000	1000	480	1000	1000	580	1000	1000	680	1000	1000	780	1000	1000
2900	1000	1000	390	1000	1000	490	1000	1000	590	1000	1000	690	1000	1000	790	1000	1000
3000	1000	1000	400	1000	1000	500	1000	1000	600	1000	1000	700	1000	1000	800	1000	1000
3100	1000	1000	410	1000	1000	510	1000	1000	610	1000	1000	710	1000	1000	810	1000	1000
3200	1000	1000	420	1000	1000	520	1000	1000	620	1000	1000	720	1000	1000	820	1000	1000
3300	1000	1000	430	1000	1000	530	1000	1000	630	1000	1000	730	1000	1000	830	1000	1000
3400	1000	1000	440	1000	1000	540	1000	1000	640	1000	1000	740	1000	1000	840	1000	1000
3500	1000	1000	450	1000	1000	550	1000	1000	650	1000	1000	750	1000	1000	850	1000	1000
3600	1000	1000	460	1000	1000	560	1000	1000	660	1000	1000	760	1000	1000	860	1000	1000
3700	1000	1000	470	1000	1000	570	1000	1000	670	1000	1000	770	1000	1000	870	1000	1000
3800	1000	1000	480	1000	1000	580	1000	1000	680	1000	1000	780	1000	1000	880	1000	1000
3900	1000	1000	490	1000	1000	590	1000	1000	690	1000	1000	790	1000	1000	890	1000	1000
4000	1000	1000	500	1000	1000	600	1000	1000	700	1000	1000	800	1000	1000	900	1000	1000

perature factors for non-hydrogen and isotropic temperature factors for hydrogen atoms) brought  $R$  to 0.038.

Finally a sequence of four cycles of refinement was carried out using a set of weights based on the mean  $\Delta F$  as a function of  $F_o$ . These weights were determined as follows (Wiesner & Lingafelter, 1966): from a plot of  $\Delta F$  versus  $F_o$ , using the structure factors at  $R=0.049$ , expressions were obtained giving  $\Delta F$  as a linear function of  $F_o$  for several ranges of  $F_o$ . Weights were then set equal to  $1/(\Delta F)^2$ . Thus:

$$\begin{aligned} \sqrt{w} &= 0.0 && \text{for unobserved reflections with } F_c < F_{\min} . \\ \sqrt{w} &= 1/(8.65 - 0.0883 F_o) && \text{for unobserved reflections with } F_c \geq F_{\min} \text{ and} \\ &&& \text{for observed reflections with } F_o \leq 60 . \\ \sqrt{w} &= 1/(4.10 - 0.0125 F_o) && \text{for observed reflections with } 60 < F_o \leq 120 . \\ \sqrt{w} &= 1/2.60 && \text{for observed reflections with } F_o > 120 . \end{aligned}$$

This sequence of cycles resulted in no change in  $R$ , and only small changes in parameters, but the weighting scheme is believed to represent a reasonable reflection of the probable accuracy of the structure factors. In the final cycle the average parameter shift was  $0.02\sigma$  and the maximum was  $0.10\sigma$  for non-hydrogen and  $0.26\sigma$  for hydrogen atoms.

The final values of the parameters and their estimated standard deviations are given in Table 1 and observed and calculated structure factors in Table 2.

#### Monoclinic form

The initial structure factor calculation was carried out with the parameters from the isomorphous nickel compound (Braun & Lingafelter, 1966), giving a value of  $R=0.119$ . Refinement was carried out in exactly the same manner as for the tetragonal form after location of the hydrogen atoms. The weighting scheme used for the final sequence of cycles was

$$\begin{aligned} \sqrt{w} &= 0.0 && \text{for unobserved reflections with } F_c < F_{\min} . \\ \sqrt{w} &= 1/(3.33 - 0.0583 F_o) && \text{for unobserved reflections with } F_c \geq F_{\min} \text{ and} \\ &&& \text{for observed reflections with } F_o \leq 28 . \\ \sqrt{w} &= 1/1.70 && \text{for observed reflections with } 28 < F_o \leq 60 . \\ \sqrt{w} &= 1/(-2.59 + 0.714 F_o) && \text{for observed reflections with } 60 < F_o \leq 95 . \\ \sqrt{w} &= 1/4.20 && \text{for observed reflections with } F_o > 95 , \end{aligned}$$

and gave a final  $R$  value of 0.052. In the final cycle the average parameter shift was  $0.06\sigma$  and the maximum was  $0.48\sigma$  for non-hydrogen and  $0.59\sigma$  for hydrogen atoms.

Table 3. *Monoclinic form*

Positional parameters ( $\times 10^4$ ) and thermal parameters ( $\times 10^2$ )

	$x/a$	$y/b$	$z/c$	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
Pd	0000	5000	5000	338 (4)	422 (4)	292 (3)	12 (7)	111 (3)	-31 (8)
O	1827 (4)	5002 (11)	5955 (5)	263 (24)	819 (39)	391 (25)	10 (57)	83 (20)	-308 (52)
N	-0310 (7)	5797 (9)	6754 (8)	349 (39)	622 (45)	312 (33)	-2 (39)	111 (31)	-16 (33)
C(1)	1876 (9)	6327 (10)	7998 (10)	421 (55)	327 (49)	368 (42)	14 (43)	120 (39)	-23 (38)
C(2)	2446 (8)	5633 (9)	7133 (9)	278 (45)	391 (43)	306 (37)	-69 (35)	34 (35)	51 (33)
C(3)	3749 (9)	5512 (9)	7531 (11)	424 (52)	356 (52)	469 (46)	-84 (36)	126 (42)	-16 (35)
C(4)	4432 (12)	6207 (14)	8735 (13)	369 (64)	691 (75)	547 (63)	-139 (57)	136 (53)	-4 (54)
C(5)	3901 (14)	6933 (16)	9563 (15)	589 (83)	789 (84)	468 (59)	-262 (66)	103 (59)	-115 (59)
C(6)	2635 (13)	7025 (13)	9230 (13)	477 (83)	620 (73)	421 (53)	-121 (55)	106 (52)	-95 (50)
C(7)	0559 (10)	6360 (11)	7778 (11)	476 (61)	497 (57)	410 (48)	28 (46)	241 (44)	-109 (43)
C(8)	-1600 (10)	5846 (14)	6867 (10)	431 (56)	883 (76)	339 (45)	102 (55)	219 (40)	-137 (50)
C(9)	-1662 (19)	5321 (46)	8231 (21)	489 (94)	1759 (27)	602 (91)	-11 (162)	153 (81)	173 (148)
C(10)	-2128 (16)	7295 (21)	6568 (23)	351 (80)	1098 (143)	1098 (122)	106 (88)	107 (85)	-358 (96)
C(11)	4357 (12)	4685 (15)	6641 (15)	535 (61)	640 (123)	671 (65)	62 (62)	202 (53)	33 (59)

Positional parameters ( $\times 10^3$ ) and isotropic thermal parameters

H(4)	519 (9)	614 (9)	878 (9)	3.3 (2)
H(5)	431 (12)	721 (14)	1029 (13)	7.5 (4)
H(6)	229 (9)	738 (10)	983 (9)	3.5 (2)
H(7)	034 (8)	676 (9)	860 (10)	4.6 (2)
H(8)	-196 (7)	497 (13)	604 (9)	5.6 (2)
H(91)	-251 (11)	535 (11)	832 (11)	7.9 (3)
H(92)	-132 (10)	479 (16)	860 (12)	0.2 (4)
H(93)	-110 (13)	596 (16)	907 (15)	8.9 (4)
H(101)	-284 (12)	736 (14)	670 (13)	7.8 (4)
H(102)	-159 (7)	785 (7)	713 (6)	0.2 (1)
H(103)	-172 (11)	745 (12)	565 (12)	6.7 (3)
H(111)	414 (11)	539 (13)	568 (14)	9.7 (4)
H(112)	550 (19)	466 (22)	702 (19)	17.7 (6)
H(113)	405 (8)	363 (11)	666 (9)	4.2 (2)

Table 4. Observed and calculated structure factors for the monoclinic form. Columns are  $h$ ,  $10F_o$  and  $10F_c$ . Unobserved reflections are marked with \*.

Table with columns for Miller indices (h, k, l) and structure factor values (10Fo, 10Fc). Rows represent different reflections, with observed values and calculated values provided. Unobserved reflections are marked with an asterisk.

The final values of the parameters and their estimated standard deviations are given in Table 3 and observed and calculated structure factors in Table 4.

### Discussion

The centrosymmetric molecules of bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium are similar in configuration to those of the nickel compound and the 3-ethyl-palladium compound (Braun & Lingafelter,

1966, 1967*b*), but there are a number of significant differences in details, particularly in the case of the tetragonal form. The packing of the molecules of the two forms of bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium are shown in Figs. 1 and 2.

Fig. 3 shows the bond lengths and angles between the non-hydrogen atoms for both forms. The two sets of values agree well within twice the standard deviation. Also they agree well with those reported for similar compounds (Lingafelter & Braun, 1966; Braun & Lingafelter, 1966, 1967*a, b*). The bond distances involving the hydrogen atoms are given in Table 5.

Although the bond distances and angles show very little variation, considerable differences in the deviations from planarity are observed. Table 6 lists the deviations of selected atoms from the least-squares plane defined by the benzene ring in six recently studied *N*-isopropylsalicylaldimine chelates. The significant deviations from planarity are seen to be in the same direction for all compounds but to vary markedly in magnitude between compounds. Thus the oxygen atom is in all cases displaced from the plane in the same direction as the metal atom, while C(7) and N are displaced in the opposite direction. The considerable differences between the two forms of the bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium can only be due to intermolecular forces and it therefore seems likely

Table 5. Bond distances  
between carbon and hydrogen atoms

	Tetragonal form	Monoclinic form
C(4)—H(4)	0.84 (4) Å	0.85 (10) Å
C(5)—H(5)	0.89 (5)	0.77 (12)
C(6)—H(6)	0.97 (5)	0.87 (11)
C(7)—H(7)	0.92 (4)	1.00 (10)
C(8)—H(8)	0.91 (4)	1.16 (11)
C(9)—H(91)	0.95 (7)	0.99 (13)
C(9)—H(92)	1.14 (7)	0.68 (14)
C(9)—H(93)	0.86 (7)	1.08 (14)
C(10)—H(101)	0.63 (6)	0.85 (15)
C(10)—H(102)	1.06 (8)	0.88 (7)
C(10)—H(103)	1.03 (8)	1.15 (13)
C(11)—H(111)	0.91 (5)	1.14 (13)
C(11)—H(112)	1.15 (5)	1.23 (20)
C(11)—H(113)	0.92 (4)	1.09 (10)

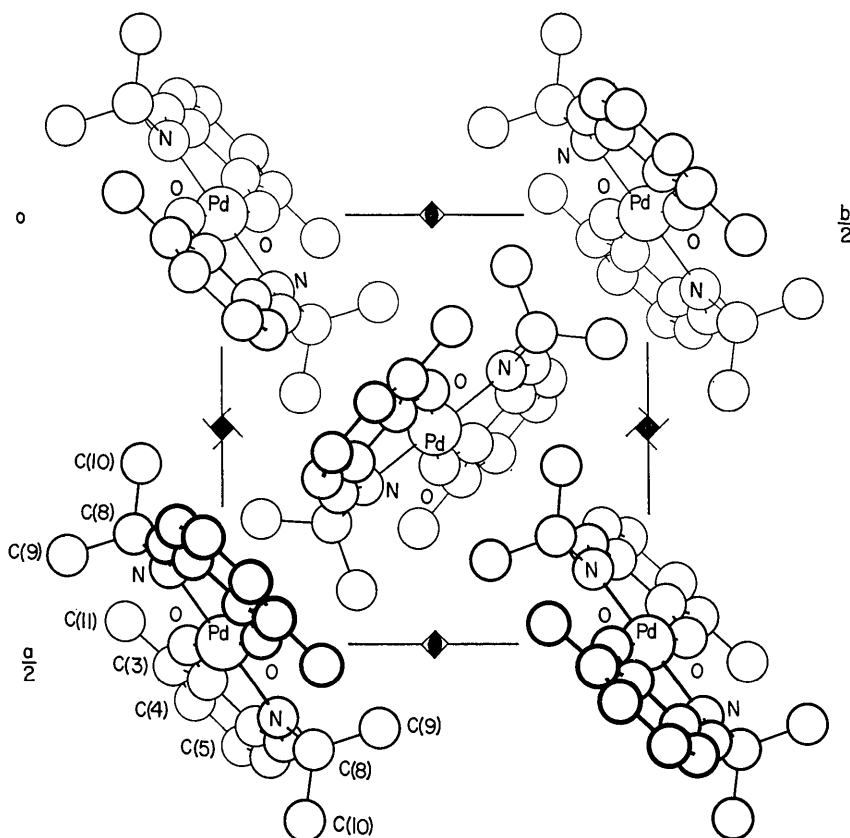


Fig. 1. Projection on (001) for the tetragonal form.

that all of the differences in deviations from planarity between different molecules may be due to inter- and intra-molecular steric effects, rather than special metal ion differences. This flexibility of these molecules has been pointed out previously (Cheeseman, Hall & Waters, 1965; Lingafelter, 1966; Boyko, Hall, Kinloch & Waters, 1966). Examination of intermolecular con-

tact distances shows none which require comment. Comparison of intra-molecular contact distances with those in the nickel compound show that the steric interactions which inhibit the formation of the planar form of bis-(*N*-isopropylsalicylaldiminato)nickel compounds are somewhat less unfavorable in the palladium chelate than in the nickel chelate.

Table 6. Distances of selected atoms from least-squares benzene plane. (Å)

Atom	(a)	(b)	(c)	(d)	(e)	(f)	(g)	(h)
Metal	-0.256	-0.212	-0.294	-0.312	-0.220	-0.210	-0.139	-0.159
O	-0.052	-0.057	-0.039	-0.066	-0.062	0.021	0.001	0.010
N	0.049	0.201	0.187	0.021	0.055	0.024	-0.021	0.047
C(7)	0.073	0.145	0.170	0.061	0.083	0.007	0.040	0.068
C(8)	0.217	0.495	0.564	0.269	0.256	-0.020	0.074	0.251
C(9)	1.427	1.695	1.930	1.693	1.633	1.036	1.351	1.730
C(10)	-1.078	-0.723	-0.536	-0.669	-0.809	-1.402	-1.112	-0.690
C(11)	0.021	0.054	0.051	0.087	-0.013	-0.013	0.008	0.007
*	0.008	0.017	0.008	0.012	0.013	0.003	0.008	0.007

- (a) Bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium, tetragonal form.  
 (b) Bis-(*N*-isopropyl-3-methylsalicylaldiminato)palladium, monoclinic form.  
 (c) Bis-(*N*-isopropyl-3-methylsalicylaldiminato)nickel (Braun & Lingafelter, 1966).  
 (d) Bis-(*N*-isopropyl-3-ethylsalicylaldiminato)nickel, ring I (Braun & Lingafelter, 1967a).  
 (e) Bis-(*N*-isopropyl-3-ethylsalicylaldiminato)nickel, ring II (Braun & Lingafelter, 1967a).  
 (f) Bis-(*N*-isopropyl-3-ethylsalicylaldiminato)palladium, (Braun & Lingafelter, 1967b).  
 (g) Bis-(*N*-isopropylsalicylaldiminato)nickel, ring I (Fox, Orioli, Lingafelter & Sacconi, 1964).  
 (h) Bis-(*N*-isopropylsalicylaldiminato)nickel, ring II (Fox, Orioli, Lingafelter & Sacconi, 1964).

\* Root-mean-square deviations for the carbon atoms of the benzene ring.

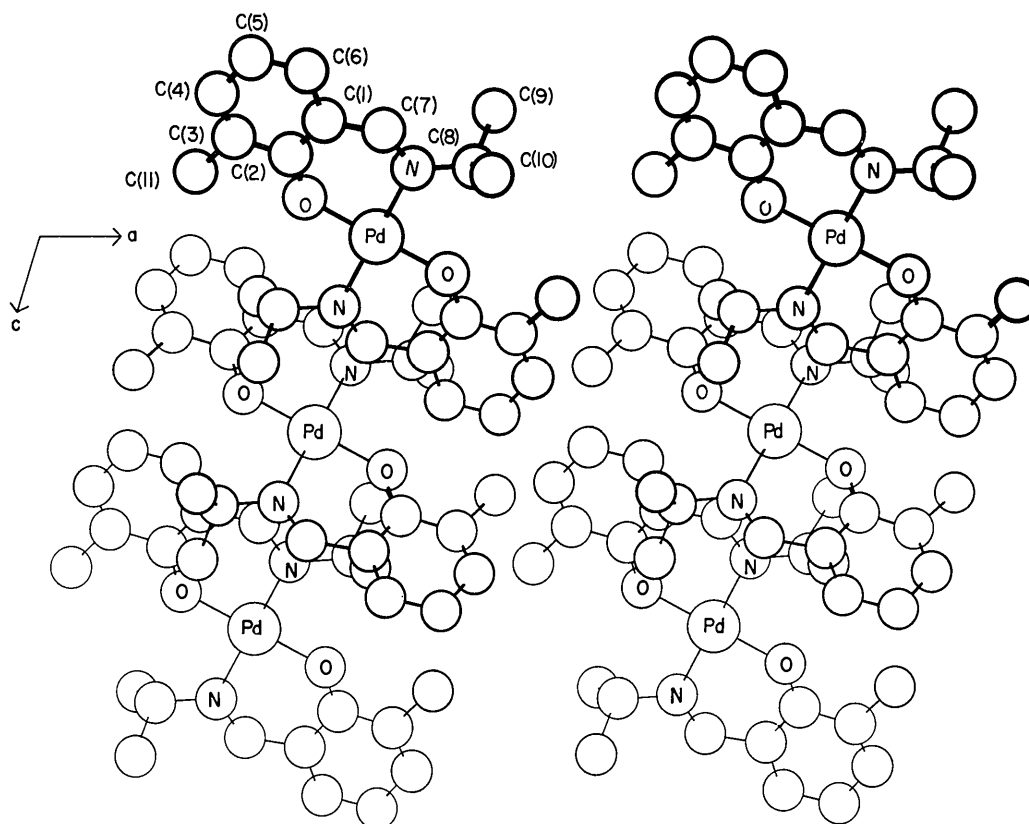


Fig. 2. Projection on (010) for the monoclinic form.

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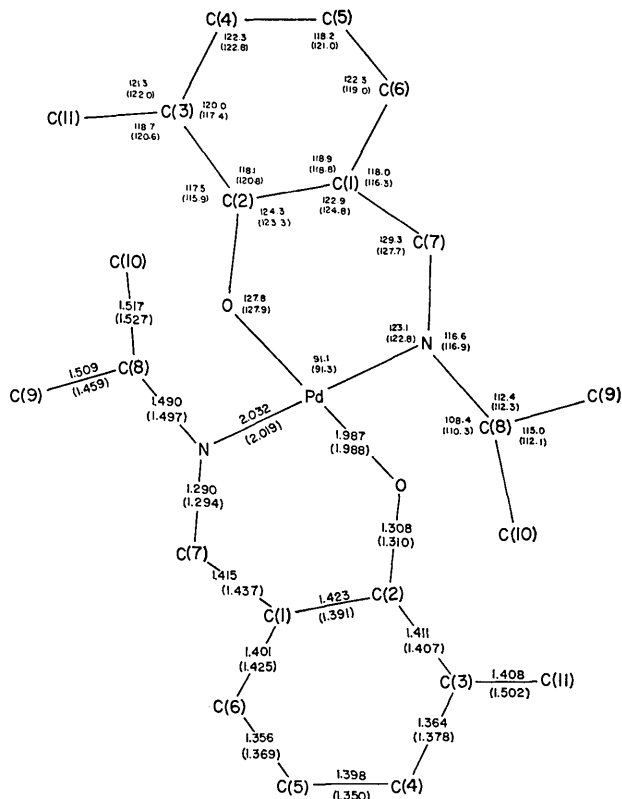


Fig. 3. Interatomic distances and angles between non-hydrogen atoms for the tetragonal and monoclinic (in parentheses) forms. Average values of the estimated standard deviations in the tetragonal form are: Pd-O, 0.003; O-C, 0.005; N-C, 0.006; C-C, 0.007; O-Pd-N, 0.1; Pd-O-C, 0.2; Pd-N-C, 0.3; N-C-C, O-C-C, C-N-C, C-C-C, 0.4. Average values of the estimated standard deviation in the monoclinic form are: Pd-O, 0.008; O-C, 0.010; N-C, 0.013; C-C, 0.018; O-Pd-N, 0.3; Pd-O-C, 0.6; Pd-N-C, 0.7; N-C-C, 1.1; O-C-C, 0.8; C-N-C, 0.9; C-C-C, 1.1.

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